

Amendments to the Specification:

Please replace the paragraph beginning at page 2, line 12 with the following amended paragraph:

Aqueous solutions are known in the art for the purpose of removing stains, corrosion or dirt from metal surfaces including, for example, Kendall, U.S. Patent No. 3,997,361, disclosing a phosphoric acid, nitric acid combination in high concentration for tarnish removal. Warner et al., U.S. Patent No. 3,640,736 and Warner, 3,846,139, teach organic thiol compositions that can be used for silver and copper tarnish removal. Combe, U.S. Patent No. 3,619,962, teaches a liquid abrasive cleaner having a substantially basic pH and an abrasive to remove tarnish from silver and copper. Kolodny et al., U.S. Patent No. 3,413,231, teach specific trivalent phosphorus compounds that impart tarnish resistance to the surface of subject metals. Similarly, Kroll et al., U.S. Patent No. 3,330,672, relies on thiol compounds, amino compounds, typically quaternary amines, and organic surfactant materials for tarnish removal. Wassell, U.S. Patent No. 2,393,866, teaches a metal tarnish remover comprising polyethylene oxide organic surfactant materials and other components in an aqueous solution for cleaning purposes. Bart, U.S. Patent No. 1,947,180, teaches that silver can be made tarnish resistant using ~~an~~ a coating of certain metals over the silver surface, typically added by an electroplating step. Other paste-like abrasive cleaning compositions are known that rely on a thickened liquid composition. Such abrasive materials, exemplified by Potter, U.S. Patent No. 4,853,000, use a thickening agent such as flour or other components with an abrasive and solvent compositions in tarnish removal. Other paste-like silver cleaners have been well known in the art for many years.

Please replace the paragraph beginning at page 3, line 12 with the following amended paragraph:

A substantial need exists in the art for an aqueous material that removes tarnish from metal surfaces rapidly and completely and also prevents the return of such tarnish for at least thirty days, preferably greater than 90 days, when exposed to normal environmental sources of sulfur and other tarnish forming compounds or uses. Of course, its activity efficacy depends entirely on the concentration of atmospheric sulfur compounds and the degree to which the metal objects are put to tarnish forming use.

Please replace the paragraph beginning at page 14, line 8 with the following amended paragraph:

Other useful salts ~~inclus de~~ include salts of the following metals. Manganese, symbol Mn has an atomic weight of 54.938049; an atomic number of 25; potential valences of 2, 4, 7; 1, 3, 5, 6 and is in Group VIIIB(7). One stable isotope occurs at the atomic weight 55. Artificial radioactive isotopes occur at the atomic weights of 49-54; 56-58. The widely-distributed, abundance averages about 0.085% of the earth's crust. Manganese occurs in the minerals pyrolusite, hausmannite, manganite, *braunite* ( $3\text{Mn}_2\text{O}_3 \cdot \text{MnSiO}_3$ ), *manganosite* ( $\text{MnO}$ ), and in several others and occurs in minute quantities in water, plants and animals. Manganese was first isolated by Gahn in 1774. The preparation of this metal is found in: John et al., cited by Mellor, *A Comprehensive Treatise on Inorganic and Theoretical Chemistry*, **12**, 163 (1932); A. H. Sully, *Manganese* (Academic Press, New York, 1955) 305 pp. A review of the physical properties of manganese is found in Meaden, *Met. Rev.* **13**, 97-114 (1968). Reviews of manganese and its compounds are found in: Kemmitt in *Comprehensive Inorganic Chemistry*, **vol. 3**, J. C. Bailar Jr. et al., Eds. (Pergamon Press, Oxford, 1973) pp 771-876; L. R. Matricardi, J. H. Downing in *Kirk-Othmer Encyclopedia of Chemical Technology* **vol. 14**, (Wiley-Interscience, New York, 3rd ed., 1981) pp 824-843. Manganese is a steel gray, lustrous, hard, brittle metal. Manganese is superficially oxidized on exposure to air, burns with an intense white light when heated in air, decreases water slowly in the cold, but rapidly on heating. Pure electrolytic manganese is not attacked by water at ordinary temperature and slightly attacked by steam. Manganese reacts with dilute mineral acids with evolution of hydrogen and formation of divalent manganous salts, reacts with aqueous solutions of sodium or potassium bicarbonate. When manganese is heated in nitrogen above  $2000^\circ$ , it burns to form a nitride. It can be converted by fluorine into di- and trifluoride and by chlorine into dichloride. In powder form, manganese reduces most metallic oxides on heating. and on heating, reacts directly with carbon, phosphorus, antimony or arsenic. Manganese has a melting point of  $1244^\circ$ ; a boiling point of  $2095^\circ$  and density of d<sub>20</sub> 7.47; d<sub>20</sub> 7.26; d<sub>1100</sub> 6.37; d<sub>1143</sub> 6.28; d<sub>20</sub> 7.21.